Additions and Corrections

Water-Soluble Receptors for Cyclic-AMP and Their Use for Evaluating Phosphate–Guanidinium Interactions [*J. Am. Chem. Soc.* **1994**, *116*, 3279–3284]. Yoko Kato, M. Morgan Conn, and Julius Rebek, Jr.*

Page 3281, last paragraph: We failed to mention that the energy enhancements seen from an additional phosphate–guanidinium interaction in our system (0.6 kcal/mol) are in good agreement with Schneider's values for exposed ionic interactions between ammonium ions and various oxyanions in water.[1] When his values are converted to 51 mM, 283 K,[2] each ionic interaction in Schneider's system contributes 0.8 kcal/mol toward binding. While we were seeking evidence for a special phosphate–guanidinium salt bridge[3] we are happy to draw attention to the work of Schneider.[4]

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(1) Schneider, H.-J. *Angew. Chem., Int. Ed. Engl.* **1989**, *28*, 753–754. Schneider, H.-J.; Blatter, T. *Angew. Chem., Int. Ed. Engl.* **1992**, *31*, 1207–1208. Schneider, H.-J.; Schiestel, T.; Zimmermann, P. *J. Am. Chem. Soc.* **1992**, *116*, 7698–7703. Schneider, H.-J.; Blatter, T.; Palm, B.; Pfingstag, U.; Rüdiger, V.; Theis, I. *J. Am. Chem. Soc.* **1992**, *114*, 7704–7708.

(2) Schneider, H.-J.; Kramer, R.; Simova, S.; Schneider, U. *J. Am. Chem. Soc.* **1988**, *110*, 6442–6448. Schneider, H.-J.; Theis, I. *J. Org. Chem.* **1992**, *57*, 3066–3070.

(3) Marquesee, S.; Baldwin, R. L. *Proc. Natl. Acad. Sci. U.S.A.* **1987**, *84*, 8898–8902.

(4) For reviews see: Schneider, H.-J. *Angew. Chem., Int. Ed. Engl.* **1991**, *30*, 1417–1436. Schneider, H.-J.; Blatter, T.; Eliseev, A.; Rüdiger, V.; Raevsky, O. *Pure Appl. Chem.* **1993**, *65*, 2329–2334.

Ab Initio Study of the Bergman Reaction: The Autoaromatization of Hex-3-ene-1,5-diyne [*J. Am. Chem. Soc.* **1994**, *116*, 4963–4969]. Roland Lindh* and B. Joakim Persson

Page 4968: In the fourth and fifth sentences we have mistakenly used "enthalpy of reaction" and "entropy of reaction" instead of "activation energy" and "differential entropy", respectively. Hence, the corrected sentences should read as follows.

In the derivation of the activation energy, Jones and Bergman estimated the differential free energy it took to "approximately half-equilibrate" the isotopic mixture. This value was then stated to be identical with the activation energy under the assumption that the differential entropy is zero.

The authors thank Prof. R. G. Bergman for calling our attention to this error.

Electrochemistry of Spontaneously Adsorbed Monolayers. Equilibrium Properties and Fundamental Electron Transfer Characteristics [*J. Am. Chem. Soc.* **1994**, *116*, 5444]. R. J. Forster and L. R. Faulkner*

Electrochemistry of Spontaneously Adsorbed Monolayers. Effects of Solvent, Potential, and Temperature on Electron Transfer Dynamics [*J. Am. Chem. Soc.* **1994**, *116*, 5453]. R. J. Forster and L. R. Faulkner*

After these papers were published, we discovered that we had inadvertently omitted reference to an important earlier paper published by D. Acevedo and H. D. Abruña [*J. Phys. Chem.* **1991**, *95*, 9590–9594], who studied adsorbed osmium–p3p complexes at Pt electrodes in aqueous solutions and other solvents. They provided isotherms for the adsorption and examined the dependence of the formal potential for the couple as the fractional coverage was varied and as the solvent was changed. They also studied the interfacial capacitance, and they established by high-speed cyclic voltammetry that the heterogeneous charge transfer kinetics are fast. We were well aware of the contribution made by Acevedo and Abruña, and we apologize for our failure to review it.

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